Suppose we have a fixed volume, $V$, containing 1 mole of $NO_2$ at pressure $p_0$.

Suppose all the

$$2NO_2(g) \rightarrow N_2(g) + 2O_2(g)$$

If $T$ remains the same, what would happen to $p$?

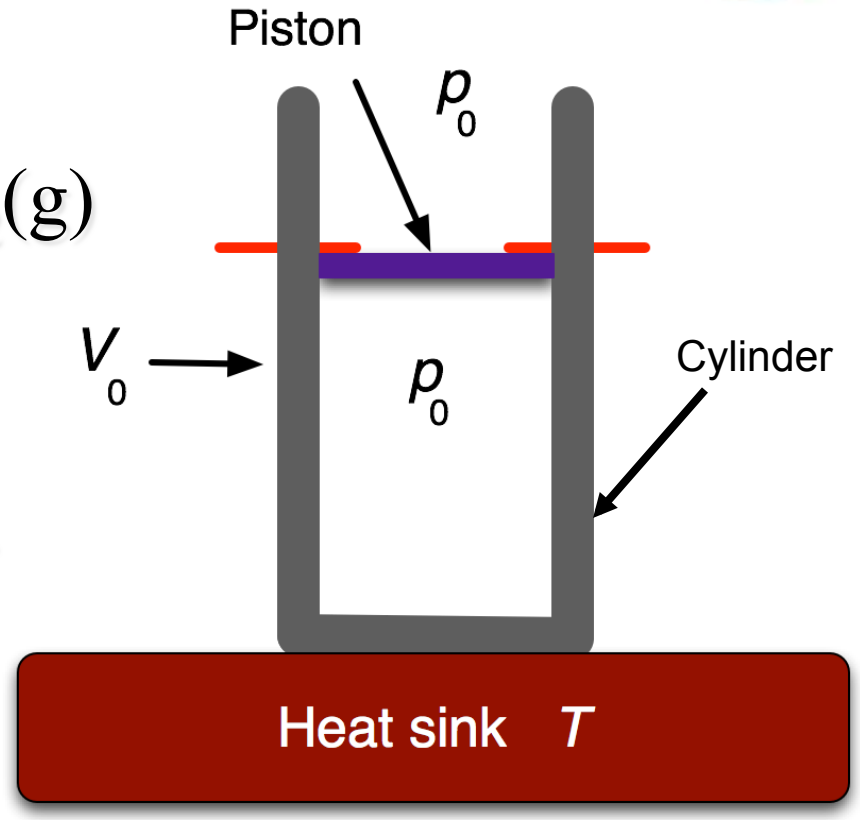
A. It would remain the same.

B. It would be 1.5 X as big.

C. It would double.

D. It would decrease but not by half.

E. It would increase but not double.



If we pull the pins holding the piston in place, the gases would expand until the pressures are equal. What would the new volume be?

A. $V_0$

B. 2/3 $V_0$

C. 3/2 $V_0$

D. Between A and B

E. Between A and C