

If I heat an enclosed volume of gas so that its Kelvin temperature doubles, what happens to the average speed of the molecules in the gas?

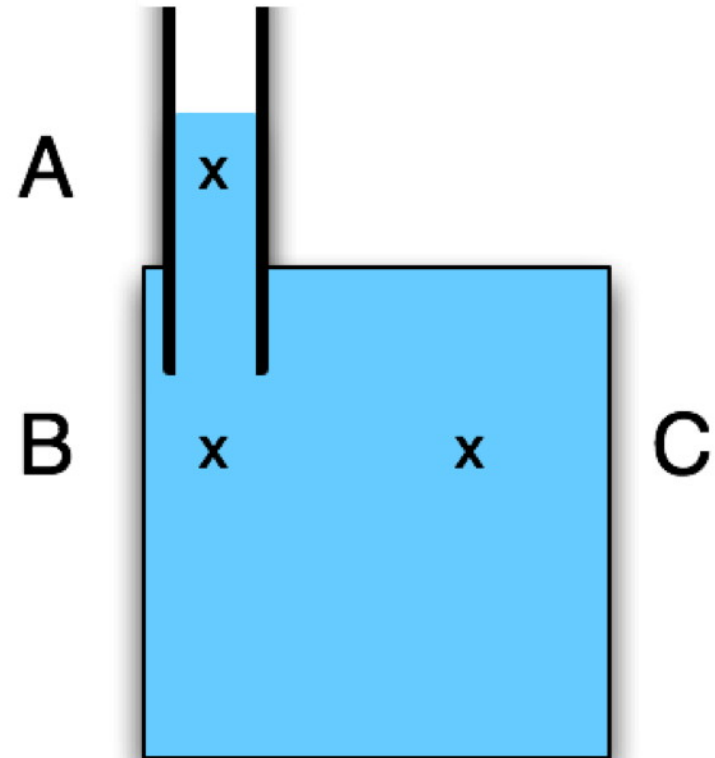


1. It more than doubles.
2. It doubles.
3. It increases by between 50% and 100%.
4. It increases but by less than 50%.
5. It stays the same
6. It decreases.



The gasoline can shown in the figure below is filled so that the gasoline goes up into the spout. How does the pressure at A and B compare?

- A. $P_A > P_B$
- B. $P_A = P_B$
- C. $P_A < P_B$
- D. You can't tell from the information given





The gasoline can shown in the figure below is filled so that the gasoline goes up into the spout. How does the pressure at B and C compare?

- A. $P_C > P_B$
- B. $P_C = P_B$
- C. $P_C < P_B$
- D. You can't tell from the information given

