For this reaction

$H_2(g) + (1/2)O_2(g) \rightarrow H_2O(g)$

$\Delta H = -286$ kJ/mol.

is $p\Delta V$

1. positive,
2. negative,
3. zero?

For this reaction (occurring in a vessel at STP)

$H_2(g) + (1/2)O_2(g) \rightarrow H_2O(g)$,

$\Delta H = -286$ kJ/mol.

is the magnitude of the internal energy change ($|\Delta U_{int}|$)

1. Greater than 286 kJ/mol?
2. Less than 286 kJ/mol?
3. Equal to 286 kJ/mol?

(Hint: Estimate the value of $p\Delta V$)