

Ammonium is NH_3 , $m_N = 14$ and $m_H = 1$ A.M.U. If 60 g of H_2 gas is available to make ammonia, how much Nitrogen is needed to use it all up?

- a) 60/14 moles = 60 g
- b) 20 moles = 280 g
- c) 30 moles = 420 g
- d) 60 moles = 840 g
- e) None is within 10%

